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Describe how the reaction coordinate can be used to predict whether a reaction will proceed or slow. Use the potential energy diagram to determine: The activation energy for the forward and reverse reactions; The difference in energy between reactants and products; The relative potential energies of the molecules at different positions on a reaction coordinate.

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**Reactions & Rates - Reaction
| Kinematics | Concentration**

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Watch a reaction proceed over time. How does total energy affect a reaction rate? Vary temperature, barrier height, and potential energies. Record concentrations and time in order to extract rate coefficients. Do temperature dependent studies to extract Arrhenius parameters. This simulation is best used with teacher guidance because it presents an analogy of chemical reactions.

Reversible Reactions - Thermodynamics | Temperature - PhET

Predict how addition of a reactant or product will

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affect the forward and reverse reaction rates, and once this new system reaches equilibrium how the reactant and product concentrations will compare to the original system at equilibrium. Calculate a rate coefficient from concentration and time data.

PhET Reactions & Rates - Rate coefficients, Activation ...

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