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GUIDELINES . Question 1 . 1. Each of
three beakers contains 25.0 mL of a 0.100
M solution of HCl, NH₃, or NH₄Cl, as
shown above.

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Answer... takes more energy to overcome the stronger hydrogen bonding in hydrazine, resulting in a higher boiling point. N_2H_4 is a polar molecule with London dispersion forces, dipole-dipole forces, and hydrogen bonding between molecules, whereas C_2H_6 is nonpolar and only has London dispersion forces.

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between molecules. It is important to the percent (c) The normal boiling point of N_2H_4 is $140^\circ C$, whereas the normal boiling point of N_2 is $-196^\circ C$. Explain, in terms of the ...

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An 8.55 mol sample of methanol, CH_3OH , is placed in a 15.0 L evacuated rigid tank and heated to 327°C .

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because it's longer than free-response sections on most other AP tests, and it includes lots of calculations and experiments that you have to interpret. The most important thing you can do is remain calm and stay focused and methodical in your approach to each question.

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from the 2011 AP Chemistry exam. The question deals with Lewis dot-diagrams, Intermolecular Forces, reactions with water...

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College Board. Visit the College Board on the Web: www.collegeboard.org. -12- 6. In an experiment, all the air in a rigid 2.0 L flask is pumped out. Then some liquid ethanol is injected into the sealed flask, which is held at 35 ° C.

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